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TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 – Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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368 Chapter 11 • Stoichiometry Section 11.1.1 Objectives

Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation.

Review Vocabulary reactant: the starting substance in a chemical reaction  
New Vocabulary stoichiometry mole ratio Defining

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## Stoichiometry

### *Chapter 11: Stoichiometry*

alanle274. Chemistry Chapter 11 Stoichiometry. stoichiometry. mole ratio. limiting reactant. excess reactant. the study of quantitative relationships between the amounts of... in a balanced equation, the ratio between the number of moles.... a reactant that is totally consumed during a chemical reaction....

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Stoichiometry Stoichiometry CHAPTER 11 SOLUTIONS

MANUAL Section 11.1 Defining Stoichiometry pages 368–372

Practice Problems pages 371–372 1. Interpret the following

balanced chemical equations in terms of particles, moles, and mass.

Show that the law of conservation of mass is

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## Stoichiometry Answer Key

What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

*Human Resources Department – Dearborn Public Schools*

In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen

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is needed to ensure complete combustion of a ...

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God, regarded as unclean the blood ...

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